

CROSS-REFERENCE TO RELATED APPLICATION

This application is a U.S. National Stage application of International Application No. PCT/US17/068100, filed on Dec. 22, 2017, which claims priority to U.S. provisional patent application No. 62/437,733, titled “Methods for Enhancing Anomalous Heat Generation,” filed on Dec. 22, 2016, which is incorporated herein in its entirety by this reference.

TECHNICAL FIELD

The present disclosure relates generally to heat generation, and more specifically, to enhancing anomalous heat generation using enriched transition metal isotopes.

BACKGROUND

For decades, scientists have been searching for alternative energy sources to replace fossil fuels and nuclear power. Over the past thirty years, scientists have, on many occasions, observed the phenomenon of excess heat being generated when hydrogen/deuterium gas is loaded into a transition metal or a metal alloy, for example, transition metals or metal alloys of palladium, nickel, or platinum. The observed excess heat generated during hydrogen/deuterium loading is often attributed to the fusion reaction between two deuterium nuclei that are trapped in the metal lattice. In one theory, two deuterium nuclei, when trapped in a metal lattice, will have a wide spread of momentum distribution based on the Heisenberg uncertainty principle. The combined probability of two deuterium nuclei having requisite momenta to overcome the Coulomb barrier may become statistically significant, triggering fusion reactions in the trapped deuterium gas. According to a second theory, the two trapped deuterium nuclei go through a quantum tunnel to reach the lower energy state, i.e., to form a ^4He nucleus.

Although these experiments have been replicated around the world, efforts to generate excess heat in a consistent manner have not been successful. Scientists have explored different conditions in which generation of excess heat can be enhanced, but research in this field has largely been inconclusive.

SUMMARY

The present disclosure relates to methods and apparatus for enhancing exothermic reactions for generating anomalous heat.

In some embodiments, an exothermic reaction between a hydrogen gas and a transition metal inside a reaction chamber is enhanced by plating the reaction chamber with an enriched product of the transition metal. The enriched product of the transition metal has an isotopic distribution that varies from the natural abundances of the stable metal isotopes. The high concentration of the isotope in the enriched product is achieved using centrifugal separation, foam fabrication, spin casting, electromagnetic calutron, laser separation, or other isotope enrichment techniques.

In one embodiment, the transition metal is palladium and one of the palladium isotopes, ^{102}Pd , ^{104}Pd , ^{105}Pd , ^{110}Pd , has a higher concentration than its natural abundance.

In one embodiment, the transition metal is nickel and one of the nickel isotopes, ^{58}Ni , ^{60}Ni , ^{61}Ni ,

^{62}Ni , and ^{64}Ni , has a higher concentration than its natural abundance.

In one embodiment, the transition metal is zirconium and one or more of the zirconium isotopes, ^{90}Zr , ^{91}Zr , ^{92}Zr , ^{94}Zr , and ^{96}Zr , have a higher concentration than its natural abundance.

In one embodiment, the transition metal is ruthenium and one or more of the ruthenium isotopes, ^{96}Ru , ^{98}Ru , ^{99}Ru , ^{100}Ru , ^{101}Ru , and ^{104}Ru , have a higher concentration than its natural abundance.

In some embodiments, the enriched product of the transition metal comprises two isotopes whose concentrations are higher than their natural abundances respectively. The two isotopes are plated on the reaction chamber in different layers. In one embodiment, the different layers of the plated isotopes are of the same geometric pattern. In another embodiment, the different layers of the plated isotopes are of different geometric patterns. The different layers of the plated isotopes may be of the same or different thicknesses.

In some embodiments, an apparatus for generating excess heat in an exothermic reaction comprises a reaction chamber and a triggering device. The reaction chamber is plated with an enriched product of a transition metal. The reaction chamber contains a hydrogen gas. The triggering device is configured to trigger the exothermic reaction between the transition metal and the hydrogen gas. The enriched product comprises an isotope of the transition metal whose concentration is higher than the natural abundance of the isotope to enhance the exothermic reaction. In some embodiments, the transition metal may be nickel or palladium.

BRIEF DESCRIPTION OF FIGURES

FIG. 1 illustrates an exemplary reactor for triggering and maintaining an exothermic reaction.

FIG. 2 is a table listing transition metal isotopes.

FIG. 3 is a chart illustrating the natural abundances of different palladium isotopes.

FIG. 4 (Tables 4A-4C) illustrates the target concentrations of various isotopes of Pd, in different embodiments.

FIG. 5 is a chart illustrating the natural abundances of different nickel isotopes.

FIG. 6 (Tables 6A-6B) illustrates the target concentrations of various isotopes of Ni, in different embodiments.

FIG. 7 is a chart illustrating the natural abundances of different zirconium isotopes.

FIG. 8 (Tables 8A-8C) illustrates the target concentrations of various isotopes of zirconium.

FIG. 9 is a chart illustrating the natural abundances of different ruthenium isotopes.

FIG. 10 (Tables 10A-10C) illustrates the target concentrations of various isotopes of ruthenium.

FIG. 11 illustrates an exemplary embodiment of plating enriched palladium or nickel to enhance anomalous heat generation in an exothermic heat generation.

FIG. 12 illustrates a second exemplary embodiment of plating enriched palladium or nickel to enhance anomalous heat generation in an exothermic heat generation.

FIG. 13 illustrates a tetragonal and cubic crystal structure.

DETAILED DESCRIPTION

FIG. 1 illustrates an exothermic reaction chamber **100**. The reaction chamber **100** comprises a metal container **102**, an electrode **104**, and a lid **106**. The interior wall of the metal container **102** is first plated with gold **108** or another material (e.g., silver). The plated gold or silver functions as a seal to prevent reaction gases in the chamber from escaping through the wall of the reaction chamber **100**. On top of the gold **108**, a layer of hydrogen absorbing material **110** is plated. Outside the reaction chamber **100**, a magnet **112** may be optionally placed.

In **FIG. 1**, the lid **106** is placed at one end of the reaction chamber **100** and is used to accommodate the electrode **104**, input/output ports **114**, and a removable electrical pass-through **116**. The electrode **104** may be made of tungsten, molybdenum, cobalt, or nickel, or other rugged metal that can withstand high voltage and high temperature environment. On the electrode **104**, a stripe of insulator **108**, such as Teflon, may be coated on the electrode **104** to prevent discharges between the electrode **104** and the exposed (i.e., un-plated) area on the interior wall of the reaction chamber **100**. The input/output ports **104** are used to introduce reaction gases into the reaction chamber **100** or extract resultant gases from the reaction chamber **100**. The input/output ports **104** can also be used to accommodate pressure controlling devices.

In some embodiments, the device used for triggering an exothermic reaction comprises a metal container and an electrode. The electrode is received through an open end of the metal container. The electrode is plated with a hydrogen absorbing material. The electrode is first plated with a layer of gold and the hydrogen absorbing material is plated on top of the layer of gold. Examples of hydrogen absorbing materials include palladium, nickel, platinum, etc.

The present disclosure teaches methods and apparatus for increasing the efficacy of an exothermic reaction by selectively enriching one or more isotopes of the hydrogen absorbing material, e.g., a transition metal such as palladium. **FIG. 2** is a table listing the naturally occurring stable isotopes of some transition metals that have large isotope distributions.

The stable isotope distributions for all transition metals are well known and documented. Although many factors—including vacancies, lattice defects, hydrogen or deuterium loading ratios, and dopants or contaminants in the Pd lattice—may be necessary for enhancing anomalous heat generation, it is believed that the isotope distribution in the Pd lattice is a critical factor in generation of anomalous heat, particularly at higher energy release levels. The methods disclosed herein relate to the deliberate and controlled modification of the isotope distribution of Pd from its natural distribution to levels necessary to generate or sustain more reliable and stronger anomalous heat generation.

In the research and tests surrounding anomalous heat generation in various physical configurations

—e.g., wet cell, gas charged tubes, and dry reactors—have been investigated. A wet cell is an electrolytic cell containing water (may be light or heavy) and an electrolyte as well as solid reactant material, wherein a voltage/current is supplied. A dry reactor is a reactor in which solid reactants can be triggered. A gas-charged tube is a reactor in which solid reactant material in a chamber can be pressurized with a gas (usually H₂ or D₂) and triggered. In any of these configurations, many have suggested that generating and sustaining reactions is specifically related to three major contributing factors: 1. defects (specifically vacancies) in the Pd metallic lattice, 2. the deuterium and hydrogen loading levels (or ratios) into the Pd lattice, and, 3. dopants (or contaminants) in the lattice structure. There has been no known discussion regarding the specific modification of the isotope distribution as a precursor or means to enhance anomalous heat generation.

Many elemental materials have a broad range of both stable and unstable (radioactive) isotopes. There are many techniques used to separate or enhance specific isotope concentrations. However, application of modified isotope concentration levels of Pd to specifically enhance the generation of anomalous heat has not been considered. There is no known background information on enhancing the ratios of specific isotopes to levels substantially higher than their natural levels to make anomalous heat generation more reliable, robust, and/or durable. Techniques for enriching particular isotopes in naturally occurring elements are widely known. The most prevalent is the enrichment of uranium 235 (radioactive uranium) using centrifuge technologies. Uranium is known to have naturally occurring abundances of U238 99.27%, U235 0.72%, and U234 of 0.0055%. Most commercial and military nuclear reactions are based on exploitation of U235 in concentrations substantially above its naturally occurring levels. These applications require uranium with concentrations of 15% to greater than 60% U235 to be viable. The most effective and prevalent technique for enriching uranium is the use of centrifuges. This technology relies on dissolving the uranium feedstock using a chemical solvent and then passing the feed material through a centrifuge system to remove higher concentration U235.

This type of technique has not been explored with respect to Pd isotope concentrations.

In existing systems and tests—including heavy water electrolytic cells, heavy water codeposition electrolytic cells, Pd lattice plated devices with hydrogen and deuterium gas, and solid state reactors—the created or used Pd lattice is based on use of commercially available solutions and/or materials. These solutions and/or materials are often palladium chloride or other solutions or solids from which a film is created using electrolysis, Physical Vapor Deposition, or Chemical Vapor Deposition. In some instances the material is an industrial Pd powder or plate. The precise isotope concentrations in the solutions or solids are not known, documented, or controlled. It is assumed that the isotope concentrations are at naturally occurring levels. However, the inconsistent efficacy of anomalous heat generation is a persistent result of these systems and tests. The exact concentrations of the Pd isotopes are controlled by specifically enriching certain isotopes above their naturally occurring levels.

The objective is to increase the efficacy of anomalous heat generation by controlling the isotope levels in the reactors by enriching specific isotopes above the naturally occurring levels. The concentration levels of ¹⁰²Pd, ¹⁰⁴Pd, ¹⁰⁵Pd, ¹¹⁰Pd to levels higher than their naturally occurring levels supports the generation of robust levels of anomalous heat, more effective control of the reaction, and enhances the durability of the reaction. Various techniques for isotope level modifications are used to achieve the enhancements of the noted isotopes to a minimum level above the natural levels. The primary technique for enhancing the concentrations is the use of a centrifuge using a Pd feedstock that is chemically dissolved.

In some embodiments, a centrifuge or equivalent mechanical techniques is used to enrich the levels of ^{102}Pd , ^{104}Pd , ^{105}Pd , ^{110}Pd to levels higher than their naturally occurring levels, while proportionally reducing the levels of ^{106}Pd and ^{108}Pd . The raw Pd stock is dissolved using a chemical solvent to break the metal at the atomic level to create feedstock. The feedstock target constituents are enriched via a centrifuge that is designed to specifically enrich Pd and not to fully separate the various isotopes from the system. **FIG. 3** shows the natural distribution of Pd. This distribution is accepted as the expected distribution of Pd in a sample of naturally occurring Pd.

In some embodiments, the concentration of one or more of the four least common isotopes is achieved to enhance the efficacy of the anomalous heat generation. By enriching one or more of the four isotopes the concentrations of the other isotopes are inherently reduced. Tables 4A-4C show the isotopes that are targeted for enrichment, and the targeted concentration/dilution of each for different embodiments.

Subsequent to enrichment the enriched Pd feedstock will be reversed into Pd metal plate, stock, or powder to support creation of the necessary materials for generation of anomalous heat.

This approach can be achieved by using a centrifuge or other technique to enrich the four least common isotopes, ^{102}Pd , ^{104}Pd , ^{106}Pd , and ^{110}Pd , which are necessary to enhance anomalous heat generation by drawing off feedstock from the centrifuge such that one or more of the four isotopes are extracted in higher concentration. It is equivalent to use a centrifuge to reduce the concentration of ^{106}Pd and ^{108}Pd . The objective is to reduce the concentration of ^{106}Pd and ^{108}Pd while enhancing the concentration of the rarer four isotopes ^{102}Pd , ^{104}Pd , ^{105}Pd and ^{110}Pd .

This same technique can be used with nickel-based anomalous heat generation. Using nickel requires the enrichment of one or more of ^{61}Ni , ^{62}Ni , and ^{64}Ni isotopes while reducing the relative concentrations of ^{58}Ni or ^{60}Ni . Nickel isotopes are shown in **FIG. 5**. Tables 6A and 6B show the isotopes that are targeted for enrichment, and the targeted concentration/dilution of each for different embodiments.

This same technique can be used with any transition-metal based anomalous heat generation, for example, zirconium or ruthenium based excess heat generation.

FIG. 7 illustrates the natural abundances of different isotopes for zirconium. Tables 8A-8C show the isotopes that are targeted for enrichment, and the targeted concentration/dilution of each for different embodiments. **FIG. 9** illustrates the natural abundances of different isotopes for ruthenium. Tables 10A-10C show the isotopes that are targeted for enrichment, and the targeted concentration/dilution of each for different embodiments.

It is also possible to generate higher concentrations of certain Pd, Ni, Zr, and Ru isotopes using rapid expansion of the metals as is common for metal foam production or spin casting. Using highly optimized foam metal fabrication techniques or spin casting techniques, specific regions of the foam or spin cast materials have higher concentrations of particular isotopes based on the relative momentum of each isotope when cast. This would require measuring the materials after fabrication to extract the region with the appropriate isotope concentrations. This is applicable to Pd, Ni, Zr, and Ru.

The enrichment of isotopes supports accentuating specific design features of the system. The following examples are the preferred embodiments for the configurations.

Table 4A: Enrich ^{105}Pd to a minimum of 30% enrichment, ^{102}Pd to a minimum of 2% enrichment, ^{104}Pd to a minimum of 13% enrichment, and ^{110}Pd to a minimum of 13% enrichment while allowing ^{106}Pd and ^{108}Pd to reduce proportionally.

Table 4B: Enrich ^{102}Pd to a minimum of 10% enrichment, ^{104}Pd to 20% enrichment, ^{105}Pd to a minimum of 25% enrichment, while allowing ^{106}Pd , ^{108}Pd , and ^{110}Pd to reduce proportionally.

Table 4C: Enrich ^{110}Pd to a minimum of 20% enrichment while allowing ^{102}Pd , ^{104}Pd , ^{105}Pd , ^{106}Pd , and ^{108}Pd to reduce proportionally.

The preferred embodiment of the Nickel Isotope Enrichment follows:

Table 6A: Enrich the ^{61}Ni to a minimum of 5% enrichment, ^{62}Ni to a minimum of 5% enrichment, and ^{64}Ni to a minimum of 5% enrichment while allowing ^{58}Ni and ^{60}Ni to reduce proportionally.

Table 6B: Enrich ^{61}Ni to a minimum of 10% enrichment while allowing ^{62}Ni , ^{64}Ni , ^{58}Ni and ^{60}Ni to reduce proportionally.

The preferred embodiments for zirconium isotopes are shown in Tables 8A-8C.

In Table 8A, ^{91}Zr is enriched to a minimum of 15% and ^{96}Zr is enriched to a minimum of 5%, while allowing ^{90}Zr , ^{92}Zr , and ^{94}Zr to be reduced proportionally,

In Table 8B, ^{91}Zr is enriched to a minimum of 15%, ^{92}Zr is enriched to a minimum of 20%, ^{94}Zr is enriched to a minimum of 20%, and ^{96}Zr is enriched to a

In Table 8C, ^{91}Zr is enriched to a minimum of 25%, while allowing ^{90}Zr , ^{92}Zr , ^{94}Zr , and ^{96}Zr to be reduced proportionally.

The preferred embodiments for ruthenium isotopes are shown in Tables 10A-10C.

In Table 10A, ^{96}Ru is enriched to a minimum of 10%, ^{98}Ru is enriched to a minimum of 5%, while allowing ^{99}Ru , ^{100}Ru , ^{101}Ru , ^{102}Ru , and ^{104}Ru to be reduced proportionally.

In Table 10B, ^{99}Ru and ^{101}Ru are enriched to a minimum of 20% each, while allowing ^{96}Ru , ^{98}Ru , ^{100}Ru , and ^{102}Ru to be reduced proportionally.

In Table 10C, ^{104}Ru is reduced to a minimum of 25%, while allowing ^{96}Ru , ^{98}Ru , ^{99}Ru , ^{100}Ru ,

^{101}Ru , and ^{102}Ru to be reduced proportionally.

Other factors to consider:

There could be many other variations of techniques for isotope enrichment for Pd, Ni, Zr, and Ru. There is the potential for foam, spin casting, or creating a plating technique where PVD or CVD technique is modified to enhance certain isotopes.

There could be enhancements to the alloy by including some of the other related materials in the alloy, e.g., palladium with rhodium and silver, and nickel with cobalt and copper.

Alloying Pd, Ni, Zr, and Ru each with higher concentrations should be considered as a method to control cost due to the abundance of various materials.

An application of these material configurations is to use the enriched isotopes, or alternatively pure isotopes, of Pd and Ni as the building block for PVD and CVD device coating. By creating Pd isotope enriched and Ni isotope enriched targets, or pure Pd or pure Ni isotope targets, for PVD and CVD individual layers of the specific isotopes.

Using pure isotopes of Pd and Ni, complex geometric structures with stratified layers of isotopes can be constructed. **FIG. 11** shows an example of a multiple isotope configuration built on a substrate.

In this embodiment a single layer of Isotope 1 is placed on the substrate material via CVD or PVD technology. Subsequently, layers of Isotopes are placed in a geometric array in a pattern on the surface. In this embodiment a configuration with cylindrical geometry isotopes stacked vertically above the base isotope that is layered on the substrate is depicted. A wide range of isotope stack geometries is feasible including cylinder, square, rectangle, and pyramid. This geometry is controlled by the sputter mask in the PVD or CVD system. The thickness of each isotope and the number of isotopes can be varied in the PVD or CVD process. The substrate material can be a rigid flat surface such that the system can be used as manufactured. Alternatively the substrate can be a very thin film such that it can be bent or shaped into various configurations. Alternatively the substrate can be a complex geometry such as a cylinder, square, or other geometry as shown in **FIG. 12**.

In this embodiment a square substrate material is coated by Isotope 1 and then place an array of square geometry stacks of various thicknesses and on each of the surfaces.

In one embodiment, the isotope structure can be oxidized under specified conditions (i.e. time, temperature, atmosphere) to create an oxide of the enriched transition metal isotope with a crystal structure different than that of the base metal. The oxide can be used as the fuel for an exothermic reaction. When exposed to hydrogen or deuterium gas, the oxide is reduced to the base metal. During the reduction, when the lattice structure changes from for example tetragonal to cubic crystal structure (see **FIG. 13**), heat and significant defects are generated. This provides a suitable environment for anomalous heat generation reactions to occur.

In another embodiment, a metal isotope with a relatively high reduction potential is oxidized under specific conditions to create an enriched oxide. This oxide is used in conjunction with either another

enriched oxide of a lower reduction potential or another enriched/non-enriched reactant, e.g. palladium or nickel. The oxide of high reduction potential can provide support if using nanoparticle reactants and/or can be a catalyst for the reaction.

The invention may be carried out in other specific ways than those herein set forth without departing from the scope and essential characteristics of the invention. The present embodiments are, therefore, to be considered in all respects as illustrative and not restrictive, and all changes coming within the meaning and equivalency range of the appended claims are intended to be embraced therein.